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Oxygen-free radicals (OFRs) arise in sequential

fashion from molecular oxygen by successive single

electron reductions. These radicals, being chemically

reactive, are capable of causing oxidative damage to

cellular components. The most common of these

radicals are superoxide anion ($O_2^{\cdot-}$), hydrogen peroxide

(H_2O_2), and hydroxyl radical ($\cdot OH$). The hydroxyl

radical is the most reactive and is capable of

causing the most extensive damage to cellular

components. The hydroxyl radical is formed

from the reaction of superoxide anion with

hydrogen peroxide in the presence of iron ions

(Fenton's reaction). The hydroxyl radical is

capable of reacting with DNA, RNA, proteins,

lipids, and other cellular components, causing

oxidative damage and cell death. The

hydroxyl radical is also capable of reacting

with other radicals, forming more stable

radicals. The hydroxyl radical is also

capable of reacting with oxygen, forming

hydroperoxyl radical ($\cdot HO_2$). The

hydroperoxyl radical is also capable of

reacting with other radicals, forming

more stable radicals. The hydroperoxyl

radical is also capable of reacting with

oxygen, forming hydroperoxide (H_2O_2).

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